

Subject	Chemistry																																																																										
	Interpretation of National Curriculum into Year group Endpoints																																																																										
Year	Term 1	Term 2	Term 3																																																																								
11	<p>Chemical Changes</p> <table border="1" data-bbox="224 458 937 1327"> <thead> <tr> <th></th> <th>Content</th> </tr> </thead> <tbody> <tr><td>1</td><td>What are metal oxides?</td></tr> <tr><td>2</td><td>What happens when metals react with other substances?</td></tr> <tr><td>3</td><td>What is a displacement reaction?</td></tr> <tr><td>4</td><td>How can metals be extracted using carbon?</td></tr> <tr><td>5</td><td>What happens to electrons in oxidation and reduction? (HT)</td></tr> <tr><td></td><td>Feedback point 1</td></tr> <tr><td>6</td><td>How do metals react with acids?</td></tr> <tr><td>7</td><td>What happens when an acid is neutralized?</td></tr> <tr><td>8</td><td>How can soluble salts be made from acids? RP</td></tr> <tr><td>9</td><td>What determines the pH and strength of an acidic or alkaline solution?</td></tr> <tr><td>10</td><td>What allows electrolysis to take place using aqueous solutions?</td></tr> <tr><td>11</td><td>How can we predict what is formed at electrodes?</td></tr> <tr><td></td><td>Feedback point 2</td></tr> <tr><td>12</td><td>How can metals be extracted using electrolysis?</td></tr> <tr><td>13</td><td>What are the products at the electrodes when an aqueous solution is electrolysed? RP</td></tr> <tr><td>14</td><td>What happens to electrons during electrolysis? (HT)</td></tr> <tr><td></td><td>Progress test</td></tr> </tbody> </table>		Content	1	What are metal oxides?	2	What happens when metals react with other substances?	3	What is a displacement reaction?	4	How can metals be extracted using carbon?	5	What happens to electrons in oxidation and reduction? (HT)		Feedback point 1	6	How do metals react with acids?	7	What happens when an acid is neutralized?	8	How can soluble salts be made from acids? RP	9	What determines the pH and strength of an acidic or alkaline solution?	10	What allows electrolysis to take place using aqueous solutions?	11	How can we predict what is formed at electrodes?		Feedback point 2	12	How can metals be extracted using electrolysis?	13	What are the products at the electrodes when an aqueous solution is electrolysed? RP	14	What happens to electrons during electrolysis? (HT)		Progress test	<p>Quantitative Chemistry</p> <table border="1" data-bbox="978 458 1691 1300"> <thead> <tr> <th></th> <th>Content</th> </tr> </thead> <tbody> <tr><td>1</td><td>Conservation of mass</td></tr> <tr><td>2</td><td>Conservation of mass & gases</td></tr> <tr><td>3</td><td>Relative formula mass & percentage mass</td></tr> <tr><td></td><td>Feedback point</td></tr> <tr><td>4</td><td>Uncertainty</td></tr> <tr><td>5</td><td>Moles</td></tr> <tr><td>6</td><td>Moles and mass</td></tr> <tr><td></td><td>Feedback point</td></tr> <tr><td>7</td><td>Stoichiometry</td></tr> <tr><td>8</td><td>Limiting reactants</td></tr> <tr><td>9</td><td>Concentration</td></tr> <tr><td>10</td><td>Yield</td></tr> <tr><td>11</td><td>Atom economy</td></tr> <tr><td>12</td><td>Titration calculations</td></tr> <tr><td>13</td><td>Gas volumes</td></tr> <tr><td></td><td>Feedback point</td></tr> <tr><td></td><td>Progress test</td></tr> </tbody> </table>		Content	1	Conservation of mass	2	Conservation of mass & gases	3	Relative formula mass & percentage mass		Feedback point	4	Uncertainty	5	Moles	6	Moles and mass		Feedback point	7	Stoichiometry	8	Limiting reactants	9	Concentration	10	Yield	11	Atom economy	12	Titration calculations	13	Gas volumes		Feedback point		Progress test	<p>Chemical cells and Fuel cells</p> <p>Students cover a very small topic on how to make batteries and how fuel cells work.</p> <p>Students will then be revising using past papers and markschemes.</p> <p>In May/June - Students will take the two exams which make up the assessment for GCSE Chemistry.</p>
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