

Subject	Chemistry					
	Interpretation of National Curriculum into Year group Endpoints					
Year	Term 1		Term 2		Term 3	
10	Chemistry of the Atmosphere		Using Resources & Equilibrium		Bonding & Structures	
		Content		Content		Content
	1	What is the atmosphere made of?	1	How do we get potable water?	1	What is ionic bonding?
	2	What was the Earth's atmosphere like in the past?	2	Can you make pure water?	2	Properties of salt
	3	How did oxygen increase and carbon dioxide decrease?	3	How do we go from pee to potable?	3	Why are alkali metals so cool?
	4	What are greenhouse gases ?	Feedback point 1		4	How can you use fire to identify elements?
	5	Why <u>is</u> climate change a problem?	4	How can we sustainably use Earth's resources	5	What do the halogens do?
	Feedback point 1		5	What is a Life Cycle Assessment?	6	How do you test for halogens?
	6	What is carbon footprint?	6	How do we reduce our impact?	Feedback point	
	Progress test		7	What causes corrosion?	7	What is special about group 0 and the transition metals?
			8	Why are alloys useful materials?	8	How do you test for metal ions?
			9	What are ceramics polymers and composites?	9	Required practical – testing for ions
			10	What are reversible reactions?	10	What is covalent bonding?
			11	What is equilibrium?	11	Why are small molecules usually gases?
			12	HT What happens when equilibrium is disturbed?	Feedback point	
		13	HT How do concentration and temperature affect equilibrium?	12	Why is diamond so strong?	
		14	HT How does pressure affect equilibrium?	13	What are polymers?	
Progress test		Progress Test		14	What is metallic bonding?	
				15	What are nanoparticles?	
				Progress test		